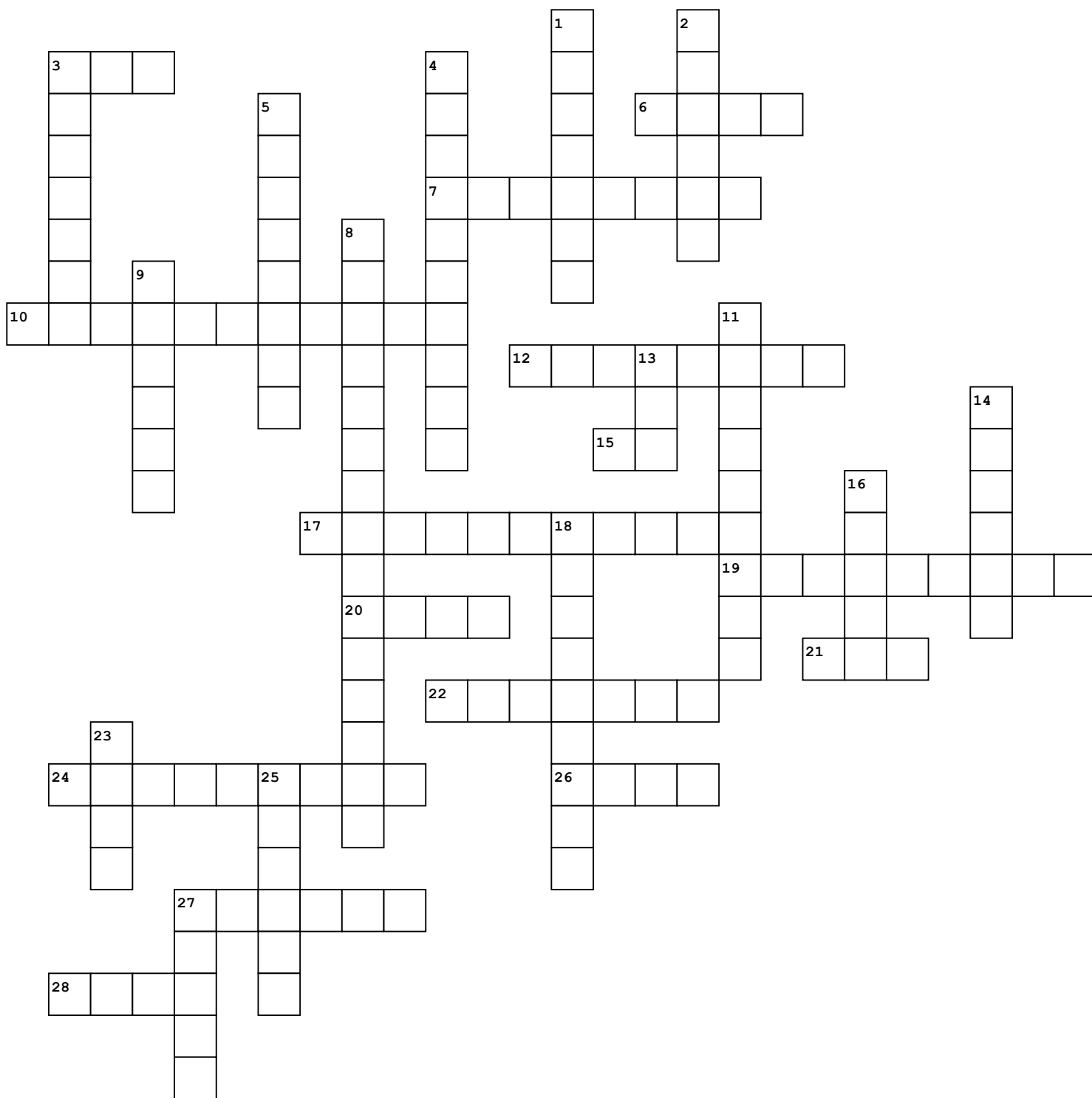


Crossword



Across

3. Each electron can only spin in ___ direction
6. For an electron to move from an energy level farther from the nucleus it must ___ energy

Down

1. ___ have the same number of electron levels and are on the same principle energy level
2. ___ are small discrete packets of energy that are emitted

7. Electrons that share the same atomic orbital have ___ spins
10. What does UV stand for?
12. There are many ___ within the Periodic Table
15. Do gas tubes filled with different gaseous element emit the same wavelengths of light?
17. The lowest potential energy arrangement of electrons in an atom is called the ___ (2w)
19. Who is the Bohr atom named after?(2w)
20. What element has the electron configuration
21. Which color corresponds to the longest wavelengths?
22. Electrons and protons ___ each other.
24. Principle energy levels are divided into _____
26. True or false, the s sub level becomes before the p sub level
27. What element has the electron configuration
28. For an electron to move from an energy closer to the nucleus it must ___ energy
3. Regions in an electron level where the probability of finding an electron is very high
4. What is the unit light is measured in?
5. Have the same number of electrons in their outer shell
8. Ground state electron configurations can be predicted by a strict set of rules known as ___(2w)
9. The wavelength of this color ranges from 590-625 (nm)
11. How many waves pass a particular point per second
13. How many electrons occupy a full 1st energy level?
14. When single electrons occupy different orbitals of the same sub level their spins are ___
16. True or false, the photon energy of all colors is the same
18. Light is ___ when it passes through a prism.
23. Electrons will occupy the next orbital only after the previous orbital is ___
25. Which color corresponds to the shortest wavelengths?
27. How fast a wave moves through space