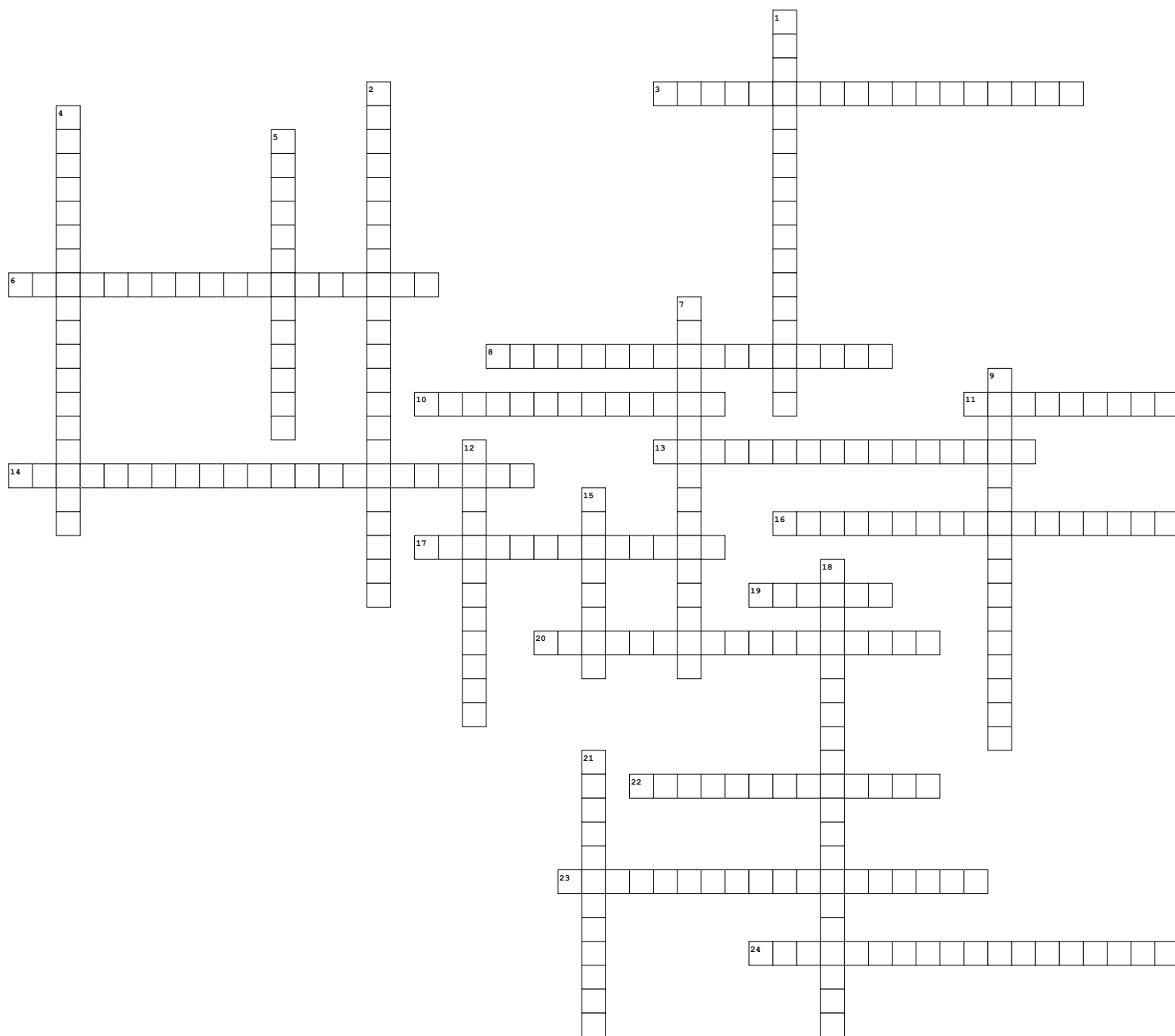


Chemistry Ch. 8



Across

3. a structure that occurs when it is possible to write two or more valid electron dot formulas that have the same number of electron pairs for a molecule or ion.
6. Two atoms held together by sharing a pair of electrons
8. The two weakest attractions between molecules
10. a tightly bound group of atoms that has a positive or negative charge and behaves as a unit.

Down

1. also known as a polar bond.
2. The energy required to break the bond between two covalently bonded atoms
4. occur when polar molecules are attracted to one another.
5. attractive forces in which a hydrogen covalently bonded to a very electronegative atom is also weakly bonded to an unshared electron pair of another electronegative atom.

11. a covalent bond between atoms in which the electrons are shared unequally.
13. the weakest of all molecular interactions, are caused by the motion of electrons.
14. a covalent bond in which one atom contributes both bonding electrons.
16. represents the covalent bonds by dashes and shows the arrangement of covalently bonded atoms.
17. one end of the molecule is slightly negative and the other end is slightly positive.
19. A molecule that has two poles is called a dipolar molecule, or
20. A compound composed of molecules.
22. solids in which all of the atoms are covalently bonded to each other.
23. A bond formed by sharing three pairs of electrons
24. A bond that involves two shared pairs of electrons
7. a molecule consisting of two atoms.
9. the chemical formula of a molecular compound.
12. The atoms held together by sharing electrons are joined by a
15. a neutral group of atoms joined together by covalent bonds.
18. When the atoms in the bond pull equally (as occurs when identical atoms are bonded), the bonding electrons are shared equally, and the bond is a
21. A pair of valence electrons that is not shared between atoms