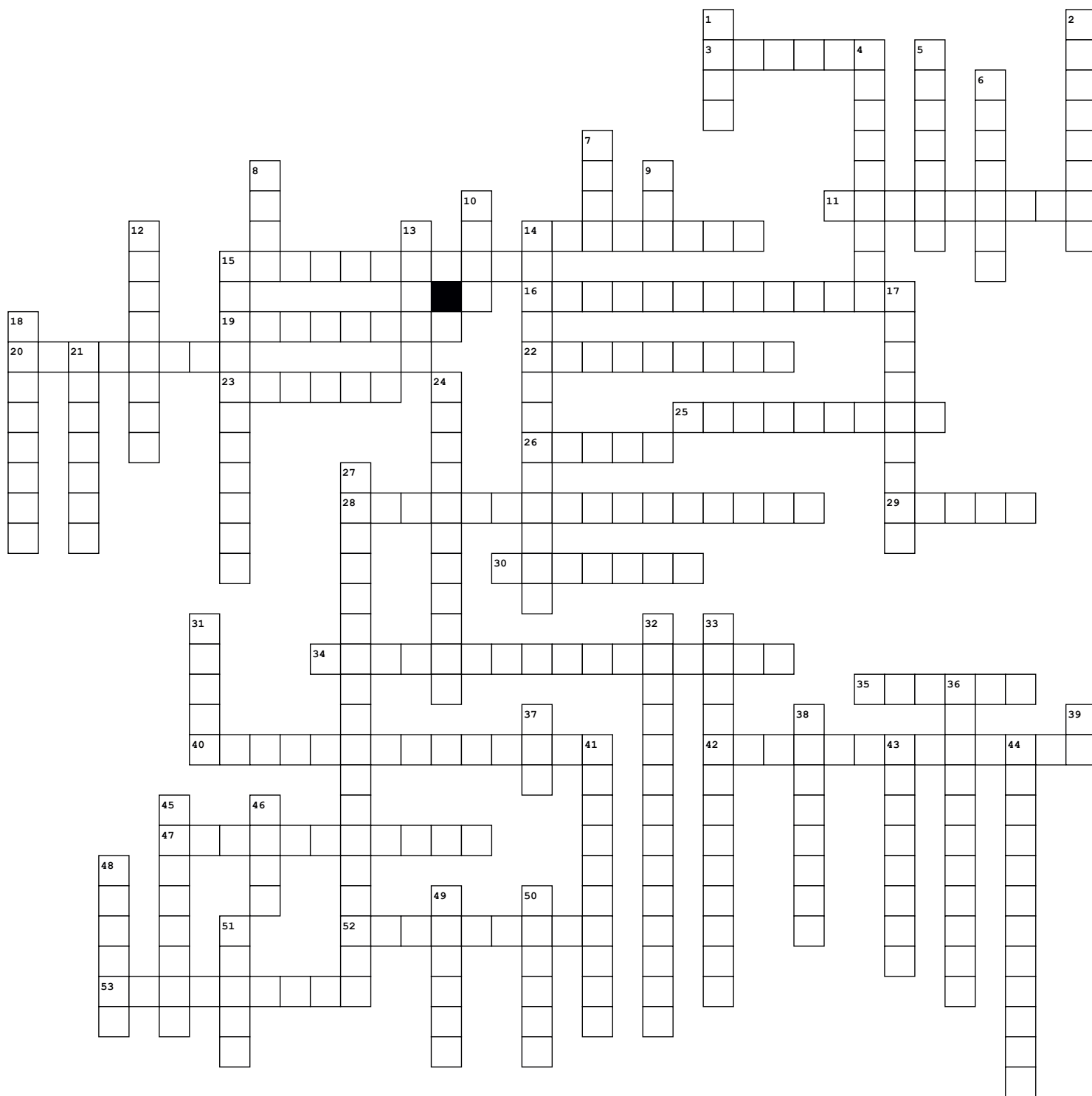


Chemistry Review



Across

3. unit of measure for volume
 11. the gaining of electrons in a reaction
 14. H, N, O, Cl, I, Br, and F all have this in common
 15. measure the average kinetic energy

Down

1. bases turn litmus paper _____
 2. these have a -1 oxidation state
 4. $A + B \rightleftharpoons AB$
 5. a substance dissolved in water

16. if HCl is an acid, Cl is its _____
19. ionic compounds _____ electrons
20. when electrons are equally distributed around a molecule
22. the mass of one mole of any element
23. metals that lose one electron
25. unshared electrons in Lewis dot structures
26. water has a pH of _____
28. the point in a reaction where all atoms are combined
29. formed when nonmetals gain electron
30. a compound that dissolves in water
34. the compound in a reaction that prevents more product from being made
35. The shape of carbon dioxide
40. the name of SO
42. elements with 2 valence electrons
47. a reaction that absorbs energy
52. slows down a chemical reaction
53. the loss of electrons in a reactions
6. energy of motion
7. covalent prefix for 6
8. the oxidation number of nitrogen in nitrate
9. the number of bonds in a molecule of oxygen
10. the oxidation number of carbon in cyanide
12. two or more elements chemically combined
13. _____ accept protons
14. $AB \rightarrow A + B$
15. the shape of carbon tetrachloride
17. the most reduced formula for a compound
18. the overall energy change in a reaction
21. ionic compounds have a(n) _____ charge
24. formed when a reaction produces an insoluble compound
27. using a more reactive metal to protect another metal
31. _____ donate protons
32. used to determine if single replacement reactions take place
33. coating nickel with a layer of zinc
36. compounds that dissociate in water and conduct electricity
37. the oxidation number of chromium in dichromate
38. speeds up a chemical reaction
39. this measure the hydrogen ion concentration
41. a reaction that has a negative enthalpy
43. elements that do not react with anything
44. transition metals require this when naming
45. these compounds create a sea of electrons
46. acids taste _____
48. formed when metals lose electrons
49. This has a definite volume or shape
50. _____ acids completely dissociate in water
51. covalent compounds _____ electrons