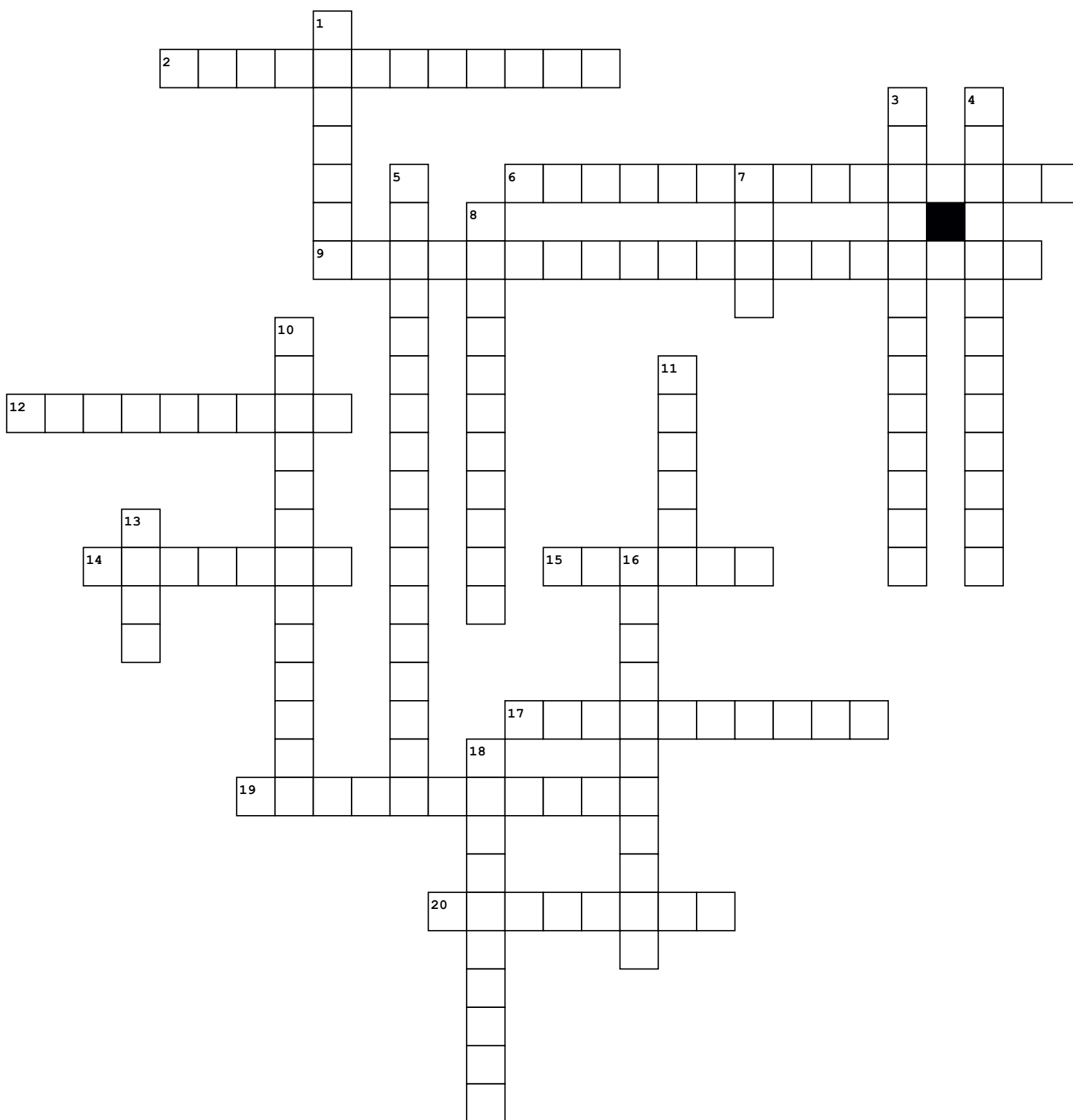


# Thermochemistry



## Across

2. the amount of energy required to raise 1 gram of a substance 1°C  
6. stored energy

## Down

1. enthalpy of +33.2 kJ/mol  
3. energy of motion  
4. enthalpy of -393.5 kJ/mol

- 9. heat change when one mole of a compound is formed in standard conditions
- 12. the enthalpy of a reaction is equal to the sum of enthalpies of individual reactions
- 14. specific heat of  $0.033 \text{ cal/g}^\circ\text{C}$
- 15. the capacity to do work or change heat
- 17. reactions that release heat
- 19. measurement of change in heat energy
- 20. most stable form of carbon
- 5. enthalpy of  $-436.7 \text{ kJ/mol}$
- 7. specific heat of  $0.46 \text{ J/g}^\circ\text{C}$
- 8. a tool used to measure changes in heat energy during reactions
- 10. most stable form of a substance at 1 atm and  $25^\circ\text{C}$
- 11. specific heat of  $0.24 \text{ J/g}^\circ\text{C}$
- 13. transfer of thermal energy ( $q$ )
- 16. reactions that require heat
- 18. enthalpy of  $-241.8 \text{ kJ/mol}$