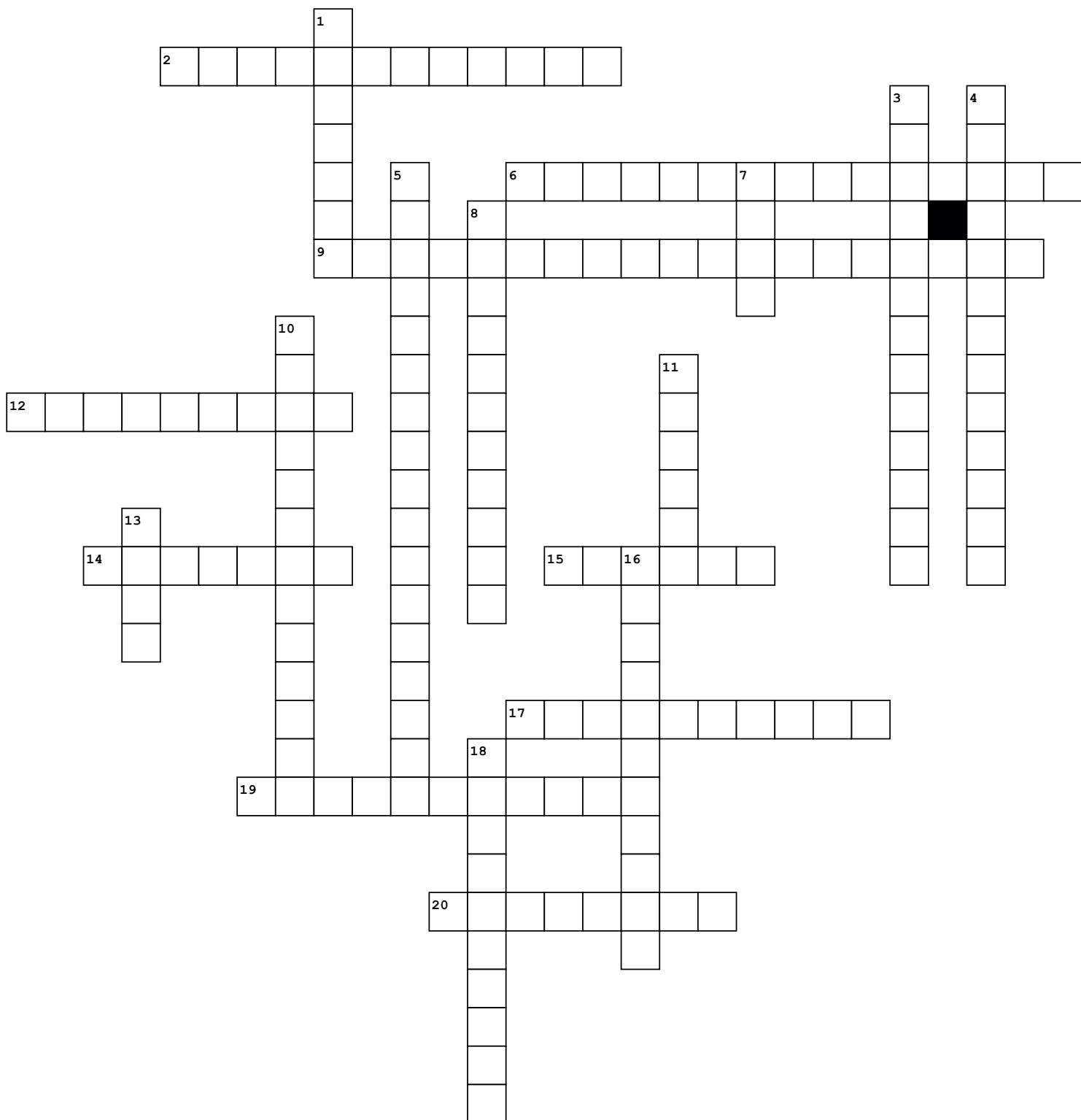


Thermochemistry



Across

2. the amount of energy required to raise 1 gram of a substance 1°C
6. stored energy

Down

1. enthalpy of +33.2 kJ/mol
3. energy of motion
4. enthalpy of -393.5 kJ/mol

- 9. heat change when one mole of a compound is formed in standard conditions
- 12. the enthalpy of a reaction is equal to the sum of enthalpies of individual reactions
- 14. specific heat of 0.033 cal/g°C
- 15. the capacity to do work or change heat
- 17. reactions that release heat
- 19. measurement of change in heat energy
- 20. most stable form of carbon
- 5. enthalpy of -436.7 kJ/mol
- 7. specific heat of 0.46 J/g°C
- 8. a tool used to measure changes in heat energy during reactions
- 10. most stable form of a substance at 1 atm and 25°C
- 11. specific heat of 0.24 J/g°C
- 13. transfer of thermal energy (q)
- 16. reactions that require heat
- 18. enthalpy of -241.8 kJ/mol